

## Empirical Formula Pogil Science Division Bellevue

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The empirical formula for a compound is C<sub>2</sub>H<sub>5</sub> and its relative formula mass is 58.0. Deduce its molecular formula. Deduce its molecular formula. (Relative atomic masses: C = 12.0, H = 1.0)

**Empirical formulae - Introducing chemical reactions - OCR ...**

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Empirical Formula Pogil Science Division Empirical formula is NaO, mass is 78 g/mole. b. Empirical formula is CH<sub>2</sub>Cl, mass is 99.0 g/mole. c. Empirical formula is C<sub>3</sub>H<sub>4</sub> mass is 121 g/ mole. stop i HSPI - The POGIL Project Limited Use by Permission Only — Not for Distribution EFMF CIYvM Extension Questions: 17. science with ms. hall - Home

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Empirical Formula Pogil Science Division Bellevue empirical formula. 2. Divide the molecular mass by the empirical mass. 3. Multiply all of the subscripts in the empirical formula by the resulting factor. Example Problem: What is the molecular formula for a material that is 92.2% C and 7.80% H by mass? The material has a molar mass of 52.0 g/mol. 1. Molecular mass /

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